

# Hydrates

**Hydrates:** are special compounds that have water groups inside of their structures. They are not chemically bonded instead they fill in holes (pockets) within the structure.

**Anhydrate:** is the compound left after water molecules have been removed.

Since the water molecules are not held in place by chemical bonds they can be easily removed by heating.

They are written with the formula

$KF \cdot xH_2O$  Where x tells how many moles of water (hydrate) are inside the compound

They are named using the following prefixes to tell how many waters there are present so  $KF \cdot 2H_2O$  is named

Potassium Fluoride Dihydrate

1	Mono	5	Penta
2	Di	6	Hexa
3	Tri	7	Hepta
4	Tetra	8	Octa

CW: # 4  $Na_2CO_3$  #5  $CaSO_4$  #6 Lead Acetate

#7Cobalt II Chloride